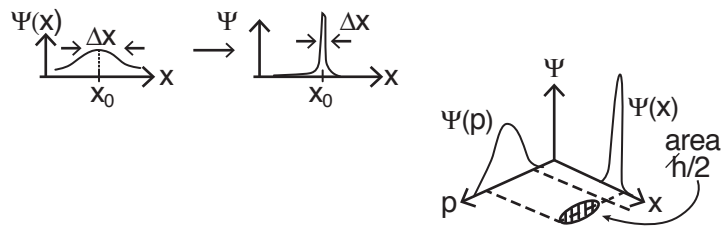


Heisenberg's (and Dirac's) view of the wavefunction:



Schrödinger initially got it wrong. He thought the wavefunction means electrons are smeared out.

However, electrons are point particles (otherwise, how could the Coulomb energy between a proton and an electron in a hydrogen atom be exactly  $-e^2/(4\pi\epsilon_0 r)$  with an exact distance  $r$ ?

Postule (4) of quantum mechanics tells us what the wavefunction REALLY is:

