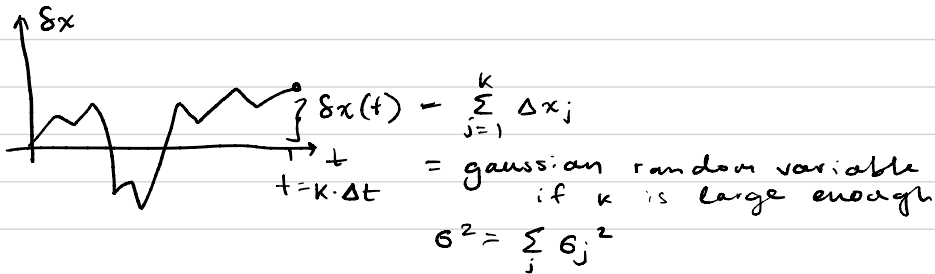


# Lecture 31

Last Time: Diffusion, force = F<sub>random</sub>



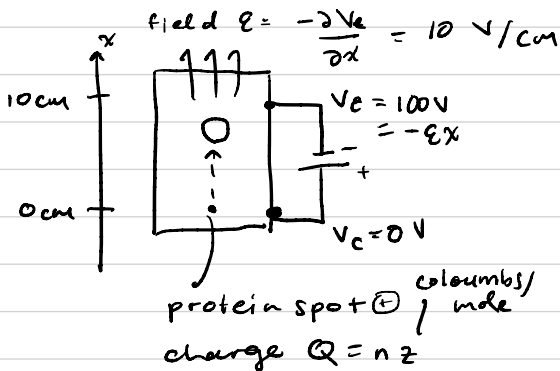
$$\langle \Delta x^2 \rangle = k \langle \Delta x_j^2 \rangle = \frac{t}{\Delta t} \langle \Delta x_j^2 \rangle$$

$$\langle \Delta x^2 \rangle = 2D \cdot t \quad \text{where } D = \frac{\langle \Delta x_j^2 \rangle}{2 \Delta t} \quad \left( \frac{m^2}{s} \right)$$

↑  
in one dimension

Today: Drift velocity and flux

ex: protein gel electrophoresis



How fast does protein move?

$$E = TS - PV + \mu n + V_e Q = TS - PV + \mu n - Ex \cdot nz$$

$$\Rightarrow \frac{\partial E}{\partial n} = \mu \Rightarrow \frac{\partial^2 E}{\partial x \partial n} = \frac{\partial \mu}{\partial x}$$

$$\Rightarrow \frac{\partial E}{\partial x} = -F_{\text{appl}} = -Enz$$

$$\frac{\partial^2 E}{\partial x \partial n} = -E \cdot z$$

But from calculus,  $\frac{\partial^2 f}{\partial x \partial y} = \frac{\partial^2 f}{\partial y \partial x}$

$$\Rightarrow -\frac{\partial \mu}{\partial x} = E z$$

$$\left( \frac{kJ}{\text{mole} \cdot \text{meter}} \right) = \left( \frac{\text{Newtons}}{\text{mole}} \right)$$

$-\frac{\partial \mu}{\partial x}$  is the driving force per mole

There are 3 different forces applied to the protein:

$$F = F_{\text{appt}} + F_{\text{random}} + F_{\text{friction}} = ma$$

friction  $\sim$  # of collisions  $\sim$  velocity

$$F = -\frac{\partial U}{\partial x} + F_{\text{random}} - \gamma \cdot v \approx 0$$

↑  
friction coefficient p. 31

Einstein's mobility  $\rightarrow u = \frac{1}{\gamma} = \frac{D}{kT}$

If we neglect (for now)  $F_{\text{random}}$ ,

$$-\frac{\partial U}{\partial x} = \gamma v$$

$$\Rightarrow qz = \gamma v \Rightarrow v_{\text{drift}} = \frac{q \cdot z}{\gamma}$$

$\gamma \rightarrow$  increases w/ field  $q$

$\rightarrow$  " " protein charge

$\rightarrow$  " " mobility ( $u = \frac{1}{\gamma}$ )

$$\gamma = 6\pi \cdot \eta \cdot r$$

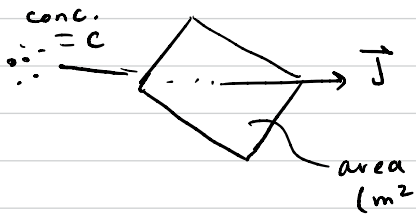
↑ viscosity      ↗ radius of protein

ex: hw 6.2  $\gamma = \frac{kT}{D}$  or  $\frac{RT}{D}$

per molecule

$R = A \cdot k$   
per mole

Flux =



$$J \left( \frac{\text{moles}}{\text{m}^2 \cdot \text{s}} \right) = v \cdot c$$

m/s      moles/m<sup>3</sup>