

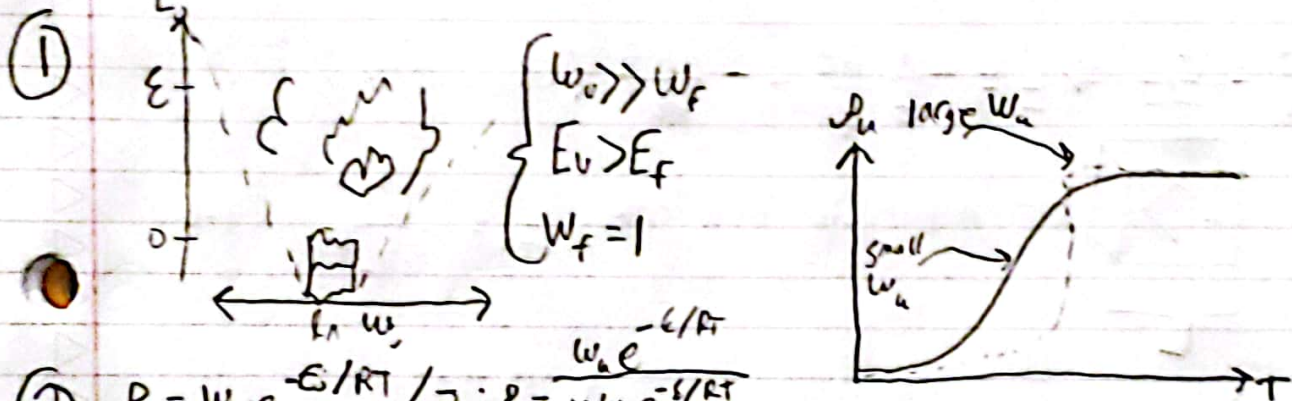
Lecture 26

Last time: $Z(T) \rightarrow \beta \rightarrow A = \sum \beta_j A_j \rightarrow$ folding!

$$Z(T) = \sum e^{-E_j/RT} = e^{-F(T)/RT}$$

Boltzmann factor
microstates at constant T
- free energy counts microstates at constant T

ex. Protein folding



② $p_j = w_j e^{-E_j/RT} / Z$; $p_u = \frac{w_u e^{-E_u/RT}}{1 + w_u e^{-E_u/RT}}$

③ $Z(T) = 1 + w_u e^{-E_u/RT}$, $E = \sum \beta_j E_j = E p_u$, $C_v = \frac{\partial E}{\partial T} \sim \frac{1}{R} \left(\frac{E^2}{T^2} \right) p_u (1-p_u)$

Today: more "laws" using $Z(T)$, $F(T)$, $G(T)$

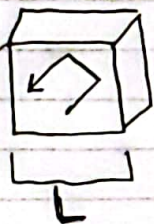
Q: What is the energy of a gas?

A: ① System + ② universal formula P_1 and P_2 + ③ combine.

① Hwt 54.2 ($\beta = \frac{1}{k_B T}$, $\frac{1}{RT}$ $R = \overset{\text{mole}^{-1}}{N_A} \cdot k_B \overset{\text{J/K}}{\text{J/K}}$)

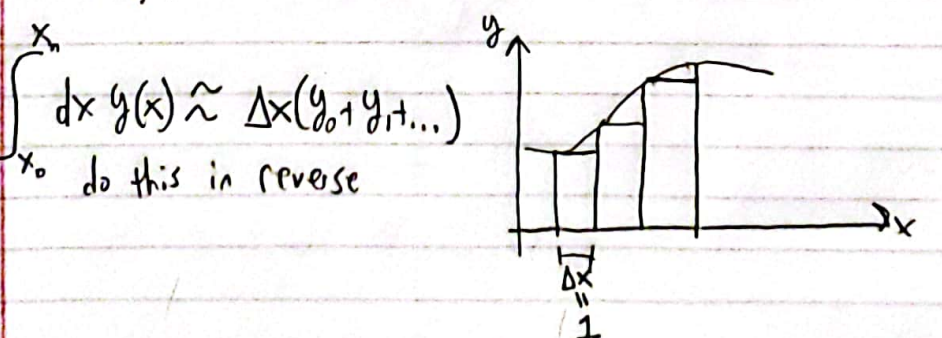
$$Z = e^{-F/kT} = e^{-\beta F} = e^{-\beta(E-TS)} = e^{-\beta E} \cdot e^{+\beta S/R}$$

$$-\frac{\partial Z}{\partial \beta} = +E \cdot Z \Rightarrow E = -\frac{1}{Z} \frac{\partial Z}{\partial \beta} = -\frac{\partial \ln Z}{\partial \beta} \quad (\text{chain rule})$$

①  For a particle in a box, $E_n = \frac{h^2 n^2}{8mL^2}$ in 1-D

In 3-D, $E_{n_x, n_y, n_z} = \frac{h^2}{8mL^2} (n_x^2 + n_y^2 + n_z^2)$ $n_x = 1, 2, \dots$
 $n_y = 1, 2, \dots$
 $n_z = 1, 2, \dots$

$Z = \sum_{n_x, n_y, n_z} e^{-\beta E_{n_x, n_y, n_z}}$; let's approximate this sum, "baby calculus in reverse"



Replacing the triple sum by triple integral,

$$Z(T) \approx \int_0^\infty dn_x \int_0^\infty dn_y \int_0^\infty dn_z e^{-\beta \frac{h^2}{8mL^2} (n_x^2 + n_y^2 + n_z^2)}$$

$$Z(T) \approx \int_0^\infty dn_x e^{-\beta \frac{h^2}{8mL^2} n_x^2} \cdot \int_0^\infty dn_y e^{-\beta \frac{h^2}{8mL^2} n_y^2} \cdot \int_0^\infty \dots$$

$$= \left\{ \int_0^\infty dn_x e^{-\beta \frac{h^2}{8mL^2} n_x^2} \right\}^3 = \left\{ \left(\frac{2\pi m}{\beta h^2} \right)^{1/2} \right\}^3 = \left(\frac{2\pi m}{\beta h^2} \right)^{3/2}$$

Final answer for E :

$$Z_{1 \text{ particle}} = \left(\frac{2\pi m}{\beta h^2} \right)^{3/2} V \quad V = L^3$$

$$Z_{N \text{ particles}} = \left(\frac{2\pi m}{\beta h^2} \right)^{3N/2} V^N$$

$$\ln Z = -\frac{3}{2} N \ln \beta + \frac{3}{2} N \ln \left(\frac{2\pi m}{h^2} \right) + N \ln V$$

$$E = -\frac{\partial \ln Z}{\partial \beta} = -\frac{\partial}{\partial \beta} \left(-\frac{3}{2} N \ln \beta \right) = \boxed{+\frac{3}{2} N k_B T = \frac{3}{2} n R T} \quad \text{Nice result!}$$